1 : Which of A - E describes a property of an ionic compound ?

A: It will dissolve in hexane.
B: It will conduct when solid.
C: It has weak forces of attraction between its molecules.
D: It will be crystalline.
E: It is usually a liquid or a gas at room temp.

2: Which of the following compounds has the greatest degree of ionic bonding?

A: Sodium chloride.
B: Potassium fluoride.
C: Caesium fluoride.
D: Lithium sulphide.
E: Magnesium bromide.

3: Which of the following is a covalent compound ?

A: Ammonia ($NH_3$)
B: Magnesium oxide.
C: Sodium chloride.
D: Barium fluoride.
E: Copper (II) sulphate.

4: Which of the following explains why molten lithium fluoride can conduct electricity?

A: Free electrons can carry electricity in the molten salt.
B: Lithium, being metallic, can conduct electricity.
C: The fluoride ions can pass an electron directly to the lithium ions in the molten salt.
D: Fluorine being non-metallic can conduct electricity.
E: The ions in lithium fluoride are free to move

5: Which of the following is a substance which can dissolve in water?

A: Sodium chloride.
B: Silicon dioxide.
C: Graphite.
D: Iodine.
E: Hydrogen.

6: Which of the following, when liquid, will release hydrogen at the positive electrode during electrolysis?

A: Lithium hydride.
B: Hydrochloric acid.
C: Water.
D: Hydrogen fluoride.
E: Methane.

7: Which of the following has the highest melting point ?

A: Silicon dioxide
B: Carbon dioxide.
C: Sulphur dioxide.
D: Nitrogen dioxide.
E: Water

8: Which of the following materials, being very hard, would make a good abrasive?

A: Glucose.
B: Sodium chloride.
C: Phosphorus chloride.
D: Silicon carbide.
E: Hydrogen bromide.

9: Which of the following is a giant molecule?

A: Ammonia.
B: Hydrogen fluoride.
C: Water.
D: Silicon dioxide.
E: Magnesium oxide.